

Ionic compounds are composed of \_\_\_\_\_ and \_\_\_\_\_. When writing an ionic formula, the total \_\_\_\_\_ charge must equal the total \_\_\_\_\_ charge because the number of electrons \_\_\_\_\_ must equal the number of electrons \_\_\_\_\_. Always write the symbol of the \_\_\_\_\_ first, followed by the symbol of the \_\_\_\_\_. The \_\_\_\_\_ method is a short-cut to correct formula writing as long as we remember to simplify our subscripts!

*In each box, write the formula of the ionic compound consisting of the positive ion to the left of the box and the negative ion above the box.*

	Cl <sup>-</sup>	S <sup>2-</sup>	F <sup>-</sup>	N <sup>3-</sup>	O <sup>2-</sup>	P <sup>3-</sup>
Mg <sup>2+</sup>						
Cs <sup>+</sup>						
Cr <sup>3+</sup>						
Na <sup>+</sup>						
Zn <sup>2+</sup>						
Al <sup>3+</sup>						